

Name:

Exit Card

Sig Figs, Density, and Percent Error

1. A sample of an element has a mass of 34.261 grams and a volume of 3.8 cubic centimeters. To which number of significant figures should the calculated density of the sample be expressed?
A) 5 B) 2 C) 3 D) 4
 2. Expressed to the correct number of significant figures, the sum of two masses is 445.2 grams. Which two masses produce this answer?
A) 210.10 g + 235.100 g
B) 210.100 g + 235.10 g
C) 210.1 g + 235.1 g
D) 210.10 g + 235.10 g
 3. What is the product of $(2.324 \text{ cm} \times 1.11 \text{ cm})$ expressed to the correct number of significant figures?
A) 2.58 cm^2 B) 2.5780 cm^2
C) 2.5796 cm^2 D) 2.57964 cm^2
 4. Which element has the greatest density at STP?
A) scandium B) selenium
C) silicon D) sodium
 5. A student measures the mass and volume of a piece of aluminum. The measurements are 25.6 grams and 9.1 cubic centimeters. The student calculates the density of the aluminum. What is the percent error of the student's calculated density of aluminum?
A) 1% B) 2% C) 3% D) 4%
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